

**Chemistry 151**  
**Worksheet 6**

Name: \_\_\_\_\_

A. Draw the Lewis dot structure for each of the following compounds.

*1. Formula: $\text{CCl}_4$ Structure:	2. Formula: $\text{HIO}_4$ Structure:
3. Formula: $\text{Na}_2\text{CO}_3$ Structure:	4. Formula: $\text{BF}_3$ Structure:

B. Classify each of the following bonds as ionic, polar covalent, or nonpolar covalent. Also, determine whether the bond is single, double, or triple.

1.  $\text{As}=\text{O}$       \_\_\_\_\_
2.  $\text{H}-\text{F}$         \_\_\_\_\_
3.  $\text{K}-\text{F}$         \_\_\_\_\_
4.  $\text{N}=\text{N}$         \_\_\_\_\_