

\* IV-2 \_\_\_\_C \_\_\_\_NC

**Chemistry 151**  
**Worksheet 7**

Name: \_\_\_\_\_

\*A. (1.2 pts.) Given the following Lewis electron dot structure, fill in the blanks.

H	Electronic geometry _____
..	Type of Hybrid _____
H : B : H	Molecular geometry _____
	Polarity _____
	No. of sigma bonds _____
	No. of pi bonds _____

B. (8.8 pts.) Draw the Lewis dot structure for each of the following and fill in the blanks.

1. SeO <sub>2</sub>	Electronic geometry _____
	Type of Hybrid _____
	Molecular geometry _____
	Polarity _____
	No. of sigma bonds _____
	No. of pi bonds _____

2. C <sub>2</sub> F <sub>2</sub>	Electronic geometry _____
	Type of Hybrid _____
	Molecular geometry _____
	Polarity _____
	No. of sigma bonds _____
	No. of pi bonds _____

3. H <sub>2</sub> Te	Electronic geometry _____
	Type of Hybrid _____
	Molecular geometry _____
	Polarity _____
	No. of sigma bonds _____
	No. of pi bonds _____

4. PI <sub>3</sub>	Electronic geometry _____
	Type of Hybrid _____
	Molecular geometry _____
	Polarity _____
	No. of sigma bonds _____
	No. of pi bonds _____